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2. Gas particles are in random, constant, straight line motion. 3. When particles collide with each other (or any container), the collisions are said

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to be elastic. 4. The volume that gas particles take up is negligible. And from the distance between particles is relatively great.

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BEHAVIOR OF GASES REVIEW

Page 102 Chemistry Unit Assessment

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11. Calculate the new temperature of a

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gas when 1500 mL at 25°C is suddenly compressed to 500 mL.

Charles' Law $K \text{ mL} \text{ K mL} \text{ V T} \text{ V T} \text{ 100}$
 $(1500) (298) (500) \text{ 1 1 2 2 12}$. A flask contains 34.6 kPa of CO₂

Student Review Packet Answer Key
Behavior Of Gases Review 2

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a measure of how fast the particles of
an object are moving 2. when it is
heated 3. Temperature of gas particles
Energy of gas particles Volume of gas
particles 1) 20°C Particles have the
smallest amount of energy. Volume is
smallest. 2) 50°C Particles have more

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introduced to deal with non-ideal
behavior of gases in Ideal gas law.

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Behavior Of Gases Review 2

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smallest. 2) 50°C Particles have more

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smallest. 2) 50°C Particles have

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a measure of how fast the particles of an object are moving 2. when it is heated 3. Temperature of gas particles
Energy of gas particles Volume of gas

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particles 1) 20°C Particles have the smallest amount of energy. Volume is smallest. 2) 50°C Particles have more energy

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You usually cannot feel it, but air has pressure. The gases in Earth's atmosphere exert pressure against everything they contact. The atmosphere rises high above Earth's surface. It contains a huge number of individual gas particles. As a result, the pressure of the tower of air above

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a given spot on Earth's surface is substantial.

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